

# Review

Acid	Neutral	Base	
$\text{H}_3\text{O}^{+1}$ $\text{H}^{+1}$ $\text{OH}^{-1}$	proton acceptor proton donor sour bitter	pH = 7 pH > 7 pH < 7	formula often ends with "OH" have "H" in 1 <sup>st</sup> part of formula alkaline

# Review

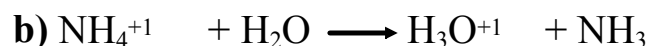
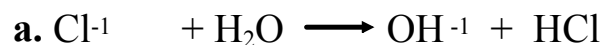
Acid	Neutral	Base
"H" in 1 <sup>st</sup> part of formula		alkaline formula often ends with "OH"
proton donor		proton acceptor
0-6.9 pH < 7	pH = 7	pH > 7 7.1 - 14
sour		bitter
H <sub>3</sub> O <sup>+</sup> H <sup>+</sup>		OH <sup>-</sup>

**1. Name the following acids and bases**

- a. KOH \_\_\_\_\_  
b. HBr \_\_\_\_\_  
c. HNO<sub>3</sub> \_\_\_\_\_  
d. H<sub>3</sub>PO<sub>3</sub> \_\_\_\_\_  
e. Be(OH)<sub>2</sub> \_\_\_\_\_  
f. H<sub>3</sub>N \_\_\_\_\_

**2. Write the formula for the following acids and bases**

- a. Carbonic acid \_\_\_\_\_  
b. Hydroiodic acid \_\_\_\_\_  
c. Nickel (III) hydroxide \_\_\_\_\_  
d. Nitric acid \_\_\_\_\_  
e. Magnesium hydroxide \_\_\_\_\_  
f. Sulfurous acid \_\_\_\_\_

**3. Identify the acid and base of the reactants.**

1. Name the following acids and bases

- a. KOH potassium hydroxide
- b. HBr hydrobromic acid
- c. HNO<sub>3</sub> nitric acid
- d. H<sub>3</sub>PO<sub>3</sub> phosphorous acid
- e. Be(OH)<sub>2</sub> beryllium hydroxide
- f. H<sub>3</sub>N hydronitric acid

2. Write the formula for the following acids and bases

- a. Carbonic acid H<sub>2</sub>CO<sub>3</sub>
- b. Hydroiodic acid HI
- c. Nickel (III) hydroxide Ni(OH)<sub>3</sub>
- d. Nitric acid HNO<sub>3</sub>
- e. Magnesium hydroxide Mg(OH)<sub>2</sub>
- f. Sulfurous acid H<sub>2</sub>SO<sub>3</sub>

3. Identify the acid and base of the reactants.

